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 **(iii) Weak acid (CH3COOH) Vs.Strong base (NaOH):** Due to poor dissociation conductivity of acid will be low.On continued addition of alkali highly ionized CH3COO-Na+ is formed and hence conductance being to increase but after neutralization further addition of alkali introduce excess of OH- ions and thus conductivity being to increase more sharply.

CH3COOH + [Na+ + OH-] =CH3COO- +Na+ +H2O

(iv) **Weak acid (CH3COOH) Vs. Weak base (NH4OH):**  In the beginning ,there will be gradual increase in the conductivity of the solution due to ionization of CH3COONH4+ which will go up to maximum near the end point after which any subsequent addition of NH4OH would not change the conductivity appreciably due to its low dissociation.

**Advantages and Disadvantages of Conductometric Titration**

Some advantages of the conductometric titration process are listed below.

* This process is very useful in the titrations of very dilute solutions and weak acids.
* The end-point of this method of titration is very sharp and accurate when compared to a few other titration processes.

The two major disadvantages of this type of titration include:

1. Only a few specific redox titrations can be done with the help of this process. This is because the conductivity of the solution is masked by relatively high hydronium ion concentration.
2. The accuracy of conductometric titration is low when the concentrations of the electrolyte are high, making the titration process unsatisfactory.